Progress Evaluation Test (PET)-1

Answer the following questions appropriately. (20*1=20)

1. Among the element given below, the one with least electronegativity is_____

   a)Lithium   b)Carbon   c)Boron   d)Flourine

2. Be, Mg,Ca, Sr and Ba are group 2 elements. Which among them will form ions most readily and why?

   _____

   _____

3. Element with Z=16 belongs to _____ period and _____ group.

4. M is metal above H in reactivity series and forms oxide M_2O. This oxide on dissolution in water forms hydroxide which is good conductor of electricity. Number of electrons in valance shell of M is_____

5. Ionization energy increases from left to right along the period because_____ and_____

6. Number of valance electrons in halogens and alkaline earth metals are _____ and _____

7. Oxidizing power of elements increases from left to right along a period. Why?

   _____

   _____

8. Which of the following has electron affinity(Electron gain enthalpy) zero?_____

   a)Ne   b)Ar   b)Na   d)S

9. Correct order of atomic size for period 3 elements is_____

10. Chlorine has higher value of electron affinity than that of Flourine due to _____

    _____

11. Metallic character _____ down the group and _____ _____along period.

12. For the species(ions) having same number of electrons, greater the _____, smaller is the_____

13. Define periodicity._____

_____

_____

14. Write electronic configuration and valency of $^{32}_{16}S$.EC:_____;Valency_____

15. Least reactive element in period 3 is_____

16. Correct match for the following is_____

A. Atomic size	1. Diagonal relationship with Li
B. Sulphur	2. Increases down the group
C. Magnesium	3. Non metal

17. Total energy required to remove outermost electron from the valance shell of an isolated gaseous atom is_____

18. State mendeleeve's periodic law._____

_____

_____

19. An element that has maximum electronegativity combines with metal M of first group. The compound is_____

20. State Modern periodic law_____

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